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1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$
- a. What is the value of the coefficient x in this equation? 4
b. What is the molar mass of C_3H_4 ? 40.07 g/mol
c. What is the mole ratio of O_2 to H_2O ? 2:1

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are produced from 7.89 moles of zinc?

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Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. _____
The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H

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CHAPTER 9 REVIEW. Stoichiometry. MIXED REVIEW. SHORT ANSWER Answer the following questions in the space provided.

1. Given the following equation: $C_3H_4(g) + x O_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. What is the value of the coefficient x in this equation? b. What is the molar mass of C_3H_4 ? c. How many moles are in an 8.0 g sample of C_3H_4 ? 2. a. What is meant by ideal conditions

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Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas:

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Stoichiometry b. Theoretically, how many moles of NH_3 will be produced? PROBLEMS Write the answer on the line to the left, Show all your work in the space provided. 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the ...

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Modern Chemistry Chapter 9 Mixed Review Stoichiometry Answers

Composition stoichiometry deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction. Reaction stoichiometry, the subject of this chapter, is based on chemical equations and the law of conservation of mass. All reaction stoichiometry

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space

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provided 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield 2 60 mol of N₂ are mixed with 120 mol of H₂ according to the following equation: N₂(g) + 3H₂(g)

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